A course syllabus is not the same as a course outline. A course syllabus outlines the general requirements for a course. A course outline is the specific document created by the individual faculty member to distribute to a specific course section. This is an “abbreviated” course syllabus because it is only collecting information on the course number, title, description, and learning outcomes. Please submit this completed form electronically to Dean Britton.

PLEASE NOTE: Any changes made to the Course Number, Title, or Catalog Description must go through the regular faculty governance process. This Expedited Process of Approval, which expires in March 2012, only pertains to approval of the Learning Outcomes. Therefore, this is NOT the form to be used to change course numbers, titles, or descriptions. This is NOT the form to use for proposing a new course. (See the Governance website for those types of proposals.)

I. Course Number and Title:
CHE100, General Chemistry

II. Catalog Description:
One-semester course for students enrolled in health careers or other curricula. Lecture topics include atomic structure, bonding, chemical equations, energy and change, gas laws, acid-base chemistry, solutions, and chemical equilibria. Prepares students for enrollment in subsequent chemistry courses. Laboratory techniques are introduced and followed by experiments which illustrate basic principles presented in lectures. (4 credit hours)

The course involves 3 hours of lecture and 3 hours of laboratory per week. Prerequisite: MAT007 or high school Sequential Math 1 or equivalent.

Notes:
1. Credit given for CHE100 or CHE 122 but not both.
2. It fulfills the SUNY General Education Requirement for Natural Sciences.

III. *Learning Outcomes:
This course is intended to prepare students enrolled in health careers or other curricula.

The successful student will demonstrate proficiency in:

- Solving problems involving atomic structure.
- Solving problems involving chemical bonding.
- Chemical nomenclature.
- Balancing and interpreting chemical equations.
- Solving problems involving chemical reactions.
- Comprehending gas-laws and chemical equilibria.
- Solving problems involving acids, bases and solutions.
- Solving stoichiometry problems.
- Performing basic laboratory techniques and experiments, which illustrate the chemical principles presented in lecture.

*These statements must appear verbatim in course outlines. However, additional outcomes may be added to individual course outlines at the instructor’s discretion.

Revised 1/10
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Revised 1/10